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### **Chapter 6 Chemical Bonding Section**

CHAPTER 6 REVIEW

Chemical Bonding

SECTION 1 SHORT

ANSWER Answer the

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Bonding Section 1.4

following questions in the space provided.

1.4 a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2. b A covalent bond consists of (a) a shared electron.

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Chapter 6 Chemical  
Bonding Bonding  
between Electroneg.  
More-neg- sulfur and  
difference Bond type  
ative atom. hydrogen  
 $2.5 - 2.1 = 0.4$  polar-  
covalent sulfur cesium  
 $2.5 - 0.7 = 1.8$  ionic  
sulfur chlorine  $3.0 - 2.5$   
 $= 0.5$  polar-covalent  
chlorine. Chemical  
Bonding, continued.  
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**Chapter 6 Chemical  
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Chemical Bonding  
section 1 Introduction  
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Geometry CHAPTER 6.  
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nonpolar-covalent  
Answers  
bond ...

## **CHAPTER 6 hemical Bonding**

Modern Chemistry 9  
Chemical Bonding  
CHAPTER 6 STUDY  
GUIDE Chemical  
Bonding SECTION 3  
IONIC BONDING AND  
IONIC COMPOUNDS  
SHORT ANSWER

Answer the following

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Answers

questions in the space provided. 1.      The notation for sodium chloride, NaCl, stands for one (a) formula unit. (c) crystal.

## **CHAPTER 6 Chemical Bonding - mchsapchemistry.com**

IONIC BONDING

COVALENT BONDING

Atoms A Atoms B Atom

C Atom D Electrons

transferred from atoms

A to atoms B Electron

pair shared between

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## 6 Chemical Bonding Section 4

atom C and atom D +  
+ Many atoms Two  
atoms Atom C Atom D  
Cation A Anion B + + +  
+ + + +----- - SECTION

6.1 Nature favors  
arrangements in which  
potential energy is  
minimized. For  
example, a boulder is  
less likely to balance

### **CHAPTER 6 Chemical Bonding**

Chapter 6: Chemical  
Bonding Section 1-  
Introduction to

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## 6 Chemical Bonding Section 4

Objectives: define chemical bond; differentiate between covalent and ionic bonding; explain why bonding occurs; use the difference in electronegativity to determine whether a bond is polar covalent, nonpolar covalent or ionic

**Ch 6 -**

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Chapter 6 - Chemical

*Page 12/26*

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Bonds. Jennie L.  
Borders. Standards.  
SPS1. Students will  
investigate our current  
understanding of the  
atom. b. Compare and  
contrast ionic and  
covalent bonds in  
terms of electron  
movement. SPS2.  
Students will explore  
the nature of matter,  
its classification and its  
system for naming  
types of matter.

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*Page 13/26*

# Where To Download Chapter 6 Chemical **Bonds**

Modern Chemistry 41 4

Chemical Bonding  
CHAPTER 6 REVIEW

Chemical Bonding  
SECTION 1 SHORT

ANSWER Answer the  
following questions in  
the space provided. 1.

\_\_\_\_\_ A chemical bond  
between atoms results  
from the attraction  
between the valence  
electrons and \_\_\_\_\_ of  
different atoms. (a)  
nuclei (c) isotopes (b)  
inner electrons (d)

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Lewis structures 2.

**CHAPTER 6 REVIEW**  
**Chemical Bonding**

A hydrogen bond is a dipole - dipole attraction between a partially positive hydrogen atom and the unshared electron pair of a strongly electronegative atom such as O, N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared, the

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hydrogen bond is a weaker attraction.

## **Chapter 6 Review: Chemical Bonding Flashcards | Quizlet**

A chemical bond between atoms results from the attraction between electrons and. A shared electron pair. A covalent bond consists of. Nonpolar covalent. If two covalently bonded atoms are identical, the bond is identified



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Worksheet  
Answers

as. Polar. A covalent bond in which there is an unequal attraction for the shared electrons is.

## **Chapter 6 Section 6-1 Review Flashcards | Quizlet**

6 Chemical Bonding.  
CHAPTER 6 REVIEW.  
Chemical Bonding.  
SECTION 2. SHORT  
ANSWERS  
Answer the following questions in the space provided. 1. Use the concept of

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## 6 Chemical Bonding Section 4

potential energy to describe how a covalent bond forms between two atoms. As the atoms involved in the formation of a covalent bond approach each other, the electron-proton attraction is stronger than the electron-electron and proton-proton repulsions.

### **6 Chemical Bonding - Somerset Canyons**

Chapter 6 Notes -

*Page 18/26*

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## 6 Chemical Bonding Section 4

Chemical Bonding  
Chemical bond - A mutual electrical attraction between the nuclei and valence electrons of different atoms that binds the atoms together 6-1  
Introduction to  
Chemical Bonding I.  
Types of Chemical  
Bonding A. Ionic  
Bonding 1.

### **Chapter 6 Notes - srvhs.org**

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## 6 Chemical Bonding Section 4

Answer Key malleable and ductile but ionic-crystalline compounds are not. The metallic bond is the same in all directions throughout the metallic structure allowing the atoms to slide past each other. This sliding is why metals are ductile and malleable.

## **Chapter 6 Chemical Bonding Section 2**

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6 Chemical  
**Covalent Answer  
Key** Bonding Section 4

CHAPTER 6 REVIEW  
Chemical Bonding  
SECTION 1 SHORT

ANSWER Answer the following questions in the space provided. 1. a A chemical bond between atoms results from the attraction between the valence electrons and of different atoms. (a) nuclei (c) isotopes (b) inner electrons (d) Lewis structures 2. b A

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covalent bond consists  
of (a) a shared  
electron.

**Section 6 1**

**Introduction To  
Chemical Bonding  
Answer Key**

CHAPTER 6 Chemical  
Bonding Modern  
Chemistry 47 Chemical  
Bonding CHAPTER 6  
REVIEW Chemical  
Bonding SECTION 4  
SHORT ANSWER

Answer the following  
questions in the space

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## 6 Chemical Bonding Section 4

provided. 1. \_\_\_\_\_ In metals, the valence electrons are considered to be (a) attached to particular positive ions. (c) immobile. (b) shared by all surrounding atoms.(d) involved in ...

### **Chapter 6 Review Chemical Bonding Answer Key**

As you read in Section 6-1, nature favors chemical bonding because most atoms

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6 Chemical  
Bonding Section 4

are at lower potential energy when bonded to other atoms than they are at as

independent particles.

In the case of covalent bond formation, this idea is illustrated by a simple example, the formation of a hydrogen-hydrogen bond.

## **CHAPTER 6 Chemical Bonding**

Chapter 6 Chemical

Bonds Section 6.2



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## 6 Chemical Bonding Section 4

Covalent Bonding  
(pages 165–169) This section discusses the formation of covalent bonds and the factors that determine whether a molecule is polar or nonpolar.

### **Chapter 6 Chemical Bonds Section 6.2 Covalent Bonding**

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